**Name…………………………………………………………… Adm. no ………....................................**

**School: ………………………………………………………….. Class…………………………………….**

**Candidate’s Sign…………………………………………..…… Date: ………………..…………………**

**233/3**

**CHEMISTRY**

**PAPER 3**

**(PRACTICAL)**

**TIME: 2 ¼ HOURS**

**SET 2**

***Kenya Certificate of Secondary Education (K.C.S.E)***

***FORM THREE***

**INSTRUCTIONS TO CANDIDATES**

* Write **your name** and **index number** in the spaces provided above
* **Sign** and **write** the date of examination in the spaces provided.
* Answer **all** questions in the spaces provided in the question paper.
* You are not allowed to start working with the apparatus for the first **15 minutes** of the 2 ¼ hours allowed for this paper. This time is to enable you to **read** the questions paper and **make sure** you have all the chemicals and apparatus that you may need.
* All working **must** be clearly shown where necessary.
* Mathematical tables and electronic calculators may be used.

**FOR EXAMINER’S USE ONLY:**

|  |  |  |
| --- | --- | --- |
| **Question** | **Maximum Score** | **Candidate’s Score** |
| 1. | 20 |  |
| 2. | 8 |  |
| 3. | 12 |  |
| **TOTAL** | **40** |  |

1. ***You are provided with:***

Metal carbonate, MCO3, solid **Q**

2M hydrochloric acid, solution **P**

Sodium hydroxide, solution **R** containing 40g per litre of solution.

You are required to determine the relative atomic mass of metal **M**

**PROCEDURE**

Measure accurately 100cm3 of solution **P** into a clean 250cm3 conical flask and add all the 4.69g of solid Q, (MCO3). Shake well and wait for effervescence to stop. Label the resulting solution as S1. Pipette 25cm3 of solution **R** into a conical flask and add 2-3 dropped of phenolphthalein indicator. Fill a burette with solution S1 and titrate against the solution **R** until the end point.

Record your results in the table below. Repeat the procedure to fill the table

Table 1

|  |  |  |  |
| --- | --- | --- | --- |
|  | I | II | III |
| Final burette reading (cm3) |  |  |  |
| Initial burette reading (cm3) |  |  |  |
| Volume of solution S1 used (cm3) |  |  |  |

Calculate

(i) Average volume of S1 used (1mk)

…………………………………………………………………………………………………………

…………………………………………………………………………………………………………

(ii) Moles of sodium hydroxide, solution **R** used (2mks)

(Na=23, O=16, H=1)

…………………………………………………………………………………………………………

…………………………………………………………………………………………………………

(iii) Moles of Hydrochloric acid, solution S1 in average volume used. (2mks)

…………………………………………………………………………………………………………

…………………………………………………………………………………………………………

(iv) Moles of Hydrochloric acid, solution S1 in 100cm3 of solution (2mks)

…………………………………………………………………………………………………………

…………………………………………………………………………………………………………

(v) Moles of hydrochloric acid in 100cm3 of the original solution **P** (2mks)

…………………………………………………………………………………………………………

…………………………………………………………………………………………………………

(vi) Moles of Hydrochloric acid, solution P, that reacted with solid Q.MCO3 (2mks)

…………………………………………………………………………………………………………

…………………………………………………………………………………………………………

(vii) Moles of MCO3 that reacted (2mks)

…………………………………………………………………………………………………………

…………………………………………………………………………………………………………

(viii) The relative formula mass of MCO3 (2mks)

…………………………………………………………………………………………………………

…………………………………………………………………………………………………………

(ix) The atomic mass of **M** (1mk)

…………………………………………………………………………………………………………

2. You are provided with solid **X**. Carry out the tests below and record your observations and inferences in the table below.

(a) Place one spatula endful of solid **X** in a test-tube and add about 10cm3 distilled water. Shake well and use for test (i) below.

(i) Test 2cm3 of the solution in the test tube with red litmus paper and blue litmus paper.

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

(ii) To 2cm3 of the solution in the test tube, add spatula endful of sodium hydrogen carbonate

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

(iii) To 2cm3 of the solution, add three drops of acidified potassium Manganate VII solution.

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

(iv) Place about 4cm3 of ethanol in a test tube and add 2 drops of concentrated sulphuric acid then add a spatula endful of solid **X**. warm the mixture carefully. Shake well and pour the mixture into 20cm3 of water in beaker.

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

3. You are provided with solid **N**. Carry out the tests and record the observations and inferences in the spaces provided

(a) Dissolve one spatula endful of solid **N** in about 10cm3 of distilled water. Divide the solution into five portions

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

(b) To the 1st portion add aqueous NaOH solution dropwise until in excess

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

(c) To the 2nd portion add dilute ammonia solution dropwise until in excess

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

(d) To the 3rd portion add three drops of dilute Sulphuric (VI) acid

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

(e)To 4th portion add 3 drops of Lead (II) nitrate solution.

|  |  |
| --- | --- |
| **Observations** | **inferences** |
| 1mk | 1mk |

(f) To 5th portion add 3 drops of Lead (II) nitrate solution and warm the mixture gently

|  |  |
| --- | --- |
| **Observations** | inferences |
| 1mk | 1mk |