**SET 10**

**CHEM P2**

**MARKING SCHEME**

1.(a) (i) Element A and B

- Both have 6 electrons to achieve an octet.

(ii) Oxide of B forms an alkaline solution that turns red litmus blue.

Oxide of D forms acidic solution, that turns blue litmus red.

(iii) E has a bigger ionic radius than the ionic radius of C.

E forms ions / ionizes by gaining electrons; which C ionizes by lose of electrons.

(iv) Formula; GH2✓1 (Rej H2G)

(v) Oxide of D is molecular with weaker vander waals forces, while the oxide of B is a giant ionic structure with stronger ionic bonds.

(vi) GCO3(S) GO(S) +CO2(g) ✓1

(b) B + Cl2 BCl2

1.5 litres of Cl2 5.9375 of BCl2

24 litres of Cl2 = (5.9375 × )g BCl2

= 95g

RFM of BCl2 = 95

RAM of BCl2 = 95-71=24

Or

B + Cl2 BCl2

Moles of Cl2 used = = 0.0625 moles

0.0625 moles Cl2 = 5.9375g BCl2

1 mole = 

RAM of B = 95 – 71 =24. (a) Gas A – Carbon (iv) oxide ✓ ½

2.Gas B – Ammonia gas ✓½

(b) Liquid C – Ammonium Chloride Solution ✓ ½

Solid D - Sodium Hydrogen Carbonate ✓ ½

(c) NH4HCO3 (aq) + NaCl (aq)  NaHCO3 (s) + NH4Cl (aq)

Ca (OH) 2(aq) + 2NH4Cl (aq)  CaCl2 (aq) ++ 2NH3 (g) + 2H2O (l)

Penalize ½✓ if not balance

½✓ if there are no states

(d) Ammonia - Manufacture of fertilizers

- Manufacture of Nitric acid

- Refrigerant

- Softening water

CaCl2  - Drying agent Name ✓1

Use ✓1

(e) - Making of glass

- Softening water

- Making sodium silicate used in making detergents any two✓1

- Paper Industry

3. (i) Condenser

(ii) To indicate when a liquid is boiling, a thermometer reads a constant temperature

(iii) A

(iv) Ethanol

**🗸**

**🗸**

Reason:- It has a lower boiling of 78oC compared to water with a boiling point of 100oC

or - The liquid with the lower boiling point boils first and its vapours are condensed

**🗸**

and the condenser to be collected as the first distillate

**🗸**

(v) Fractional distillation

(vi) - To separate components of crude oil

* To isolate O2 and N2 from air
* To manufacture spirits

**🗸**

(vii)- They are immiscible liquids

* They have different but close boiling points

**🗸**

4.a) To remove any magnesium oxide coating from the surface of magnesium// To remove any

oxide film on it

b) White solid which is magnesium oxide

c) Increase in mass was due to oxygen which combined with magnesium

d) 2Mg(s) + O2(g) \_\_\_\_\_\_\_ 2MgO(s)

Penalize ½ for wrong or missing state symbols

e) The filtrate is magnesium hydroxide which is an alkaline

Red litmus paper changed blue, but blue litmus paper remained blue

II. a) N2O 1 (Nitrogen (I) oxide) – Denitrogen Oxide.

b) K2O 1 (Potassium oxide)

c) Al2O3 (Aluminium oxide)

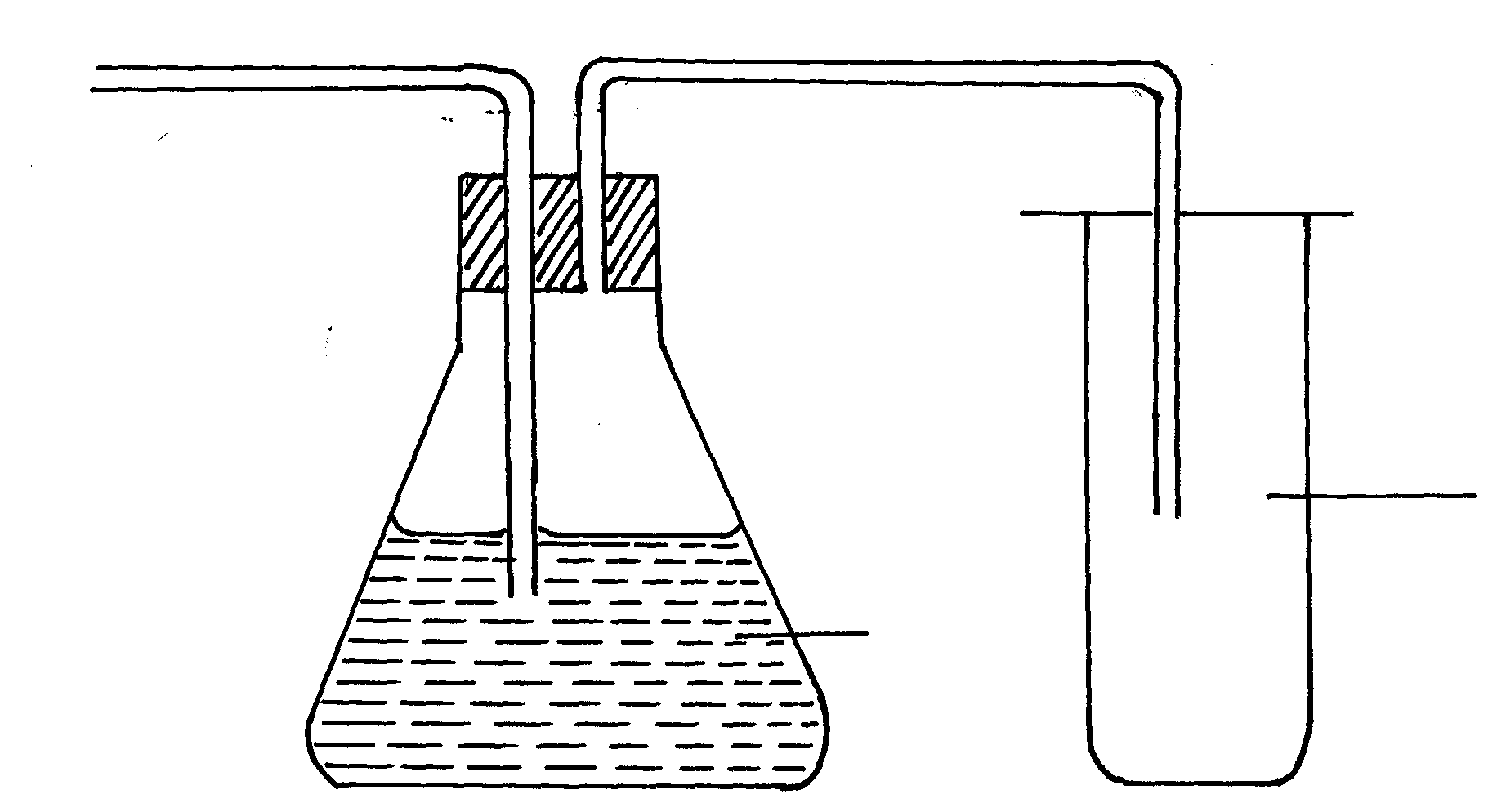
(i) Yellow lead (II) oxide turned to red then grey.

(ii) I. H2(g) + PbO(s)  H2O(l) + Pb(s)

II. 2H2(g) + O2(g)  2H2O(l)

(iii) Reducing properties of hydrogen

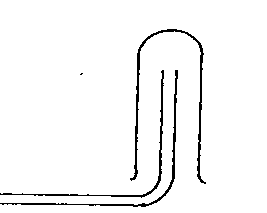
Combustion nature of hydrogen



Dry HCl

Conc. H2SO4

HCl(g)

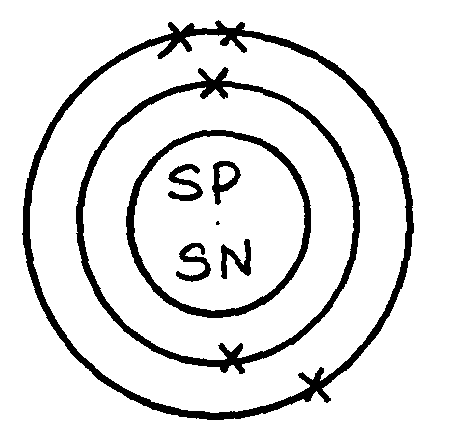
5. (a)

(b) M is hydrogen

(c) Conc. H2SO4 is a less volatile hence displaces a more volatile and from its salts i.e

(a) 5 electrons

(b) 11-5 = 6 neutrons

(c) 20/100 x 10 + 80/100 x 11 = 2+ 8.8 = 10.8

(d)

***a) 3 Mg + N2 g \_\_\_\_\_\_\_\_ Mg 3 N2 g***

***b) Argon***

***- It is inert***

***a) 3 Mg + N2 g \_\_\_\_\_\_\_\_ Mg 3 N2 g***

***b) Argon***

***- It is inert***

***c)Haber process to manufacture ammonia***

***Hydrogenation***

***Welding***

***6.a) magnesium Oxide***

***b) 2Mg(s) + O2(g) \_\_\_\_\_\_\_\_\_ 2MgO(s)***

***c) i) Sodium sulphate***

***ii) MgCO3***

***d) MgO(s) + H2SO4(aq) \_\_\_\_\_\_\_\_ MgSO4(aq) + H2O(L)***

***e) Mg2+(aq) + CO2- 3(aq)\_\_\_\_\_\_\_\_ MgCO3(s)***

***f) MgCO3(g) \_\_\_\_\_\_\_\_\_ MgO(g) + CO2(g)***

***g) Na+ ions and SO42- ions***

***h) Precipitation/ double decomposition***

7.

1. Z- Anhydrous calcium chloride ***√1mk***

Q- Water

1. Reducing agent / effect ***√1mk***

Combustible gases / burning of hydrogen in air.

1. The flame should be blown out ***√½ mk*** first as the supply of hydrogen continues to avoid explosion.√ ***½***

Heating of CuO should be ***√½ mk*** stopped to prevent re-oxidation ***√½ mk*** of hot copper before ***√½ mk*** the supply of hydrogen is stopped.

1. Hydrogen so produced is at once oxidized to water ***√1mk*** ( strong oxidizing agent )

Likelyhood of producing poisonous gases such as nitrogen (IV) oxide. ***√1mk***

1. Water molecules has lone pairs***√1mk*** of electrons which can be donated ***√mk*** and be shared with H+ to form H3O+
2. Is less dense than air / lighter than air. ***√1mk***